

Best Prictice-1

1. Title of the practice: English Language Laboratory

2. Objective of the practice:

The purpose of the language laboratory is to involve U.G. & P.G. students to actively participate in language learning exercise. The main objectives of the language laboratory are to equip the students with good communication and pronunciation skill. It also helps to prepare students for interviews and speak with correct pronunciation. It is very useful to develop all the language learning skills like listening, reading, writing and speaking. It develops not only communication skill but also develops overall personality of students. To prepare students with logical reasoning is also important objective of language laboratory, to develop resume writing techniques and sample cover letter among the students.

3. The context:

The language laboratory is becoming highly valued because it offers students a structured e-learning environment that is successful and reliable. The curriculum of the present educational system in India does not have a laboratory session for humanities and arts subjects. Only those who study science subjects they have practical work which is undertaken in a laboratory. A language laboratory for language learning is something new to Indian students where as it is very common in western countries. It provides students with the technical tools to get the best samples of pronunciation of the language.

Most of the students belong to the rural area where English is not in practice for the routine communication. Students from this area must not be lagging behind, so language laboratory is the need of time. It is helpful the make the students competent in communication.

4. The practice:

Department of English purchased the worlds most recognized 'Orell Talk' software. Teachers were trained by expert in online mode. Student accounts are opened on the server and user ID and Password is given to them. All students are permitted to join language lab according to their suitable time. Our language laboratory comprises seventeen Computers with updated Orel Talk Software. It consists of eight level of learning such as Beginners, Elementary, Pre- Intermediate, Advanced proficient and

Expert. This software also consists of near about twenty four modules. Through these modules students can develop speaking and learning skills.

The students can open their log in accessing with the password and access to the course. There are certain audio and videos to be listened and the exercises to be solved. Even grammar and vocabulary explanations are available. Student's login account can be accessed and assessed by the tutor (teacher). Student finds it easy to operate and spend their time in this useful activity. The software also records the pronunciations and answer given to the questions. It improves student's pronunciation. So, this is one of the unique and useful activities of the department. Seventeen computers are connected with LAN and each computer has a separate head phone. So, the student can carry on his learning activity without disturbing the co-learner. Student use the language laboratory in their free time. The record of their attendance is kept in register as well as their accounts are displayed on the monitor. First of all we instruct the new laboratory users how to use and handle the computers. The working of software and the courses are informed to the new students. There are some challenges and constraints in running this activity but, it is necessary for the overall development of the student. It improves their knowledge and communication in English.

Limitations:

The language laboratory would not let the English teaching learning process be effective if there are some troubles with the technology of it. Language lab software is sometimes costly and not affordable by every college. If student do not have sufficient knowledge of computers then it is useless for them. By using language lab, students cannot analyse their performance critically. In reality, student hardly find the time to engage in the lab apart from the scheduled hours, in the course of study.

5) Evidence of Success:

- ❖ Reduces the distraction in the class.
- ❖ Increases students' excitement and interest towards language learning.
- ❖ It helps teachers to deliver individualized teaching.
- ❖ It helps student to focus on assignment.
- ❖ It is easy to contact and talk with every one in the class.

- ❖ It creates equal opportunities for all to see and hear.
- ❖ language laboratory creates a change into the normal class routine.

6) problems encountered and recourses required:

The problems encountered by the teacher:

- ❖ Too long material make the student board , the teacher have difficulties to select the materials
- ❖ Teachers are unable to identify the error of the machine.
- ❖ Reluctance to take the time to introduce an internet based activity.

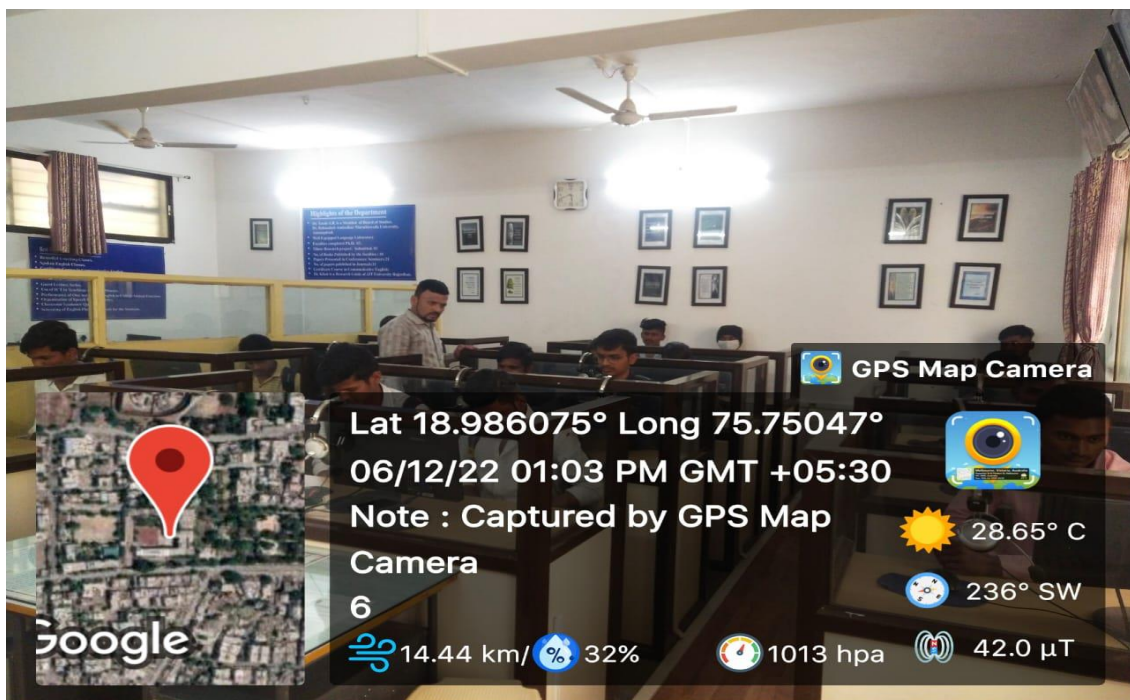
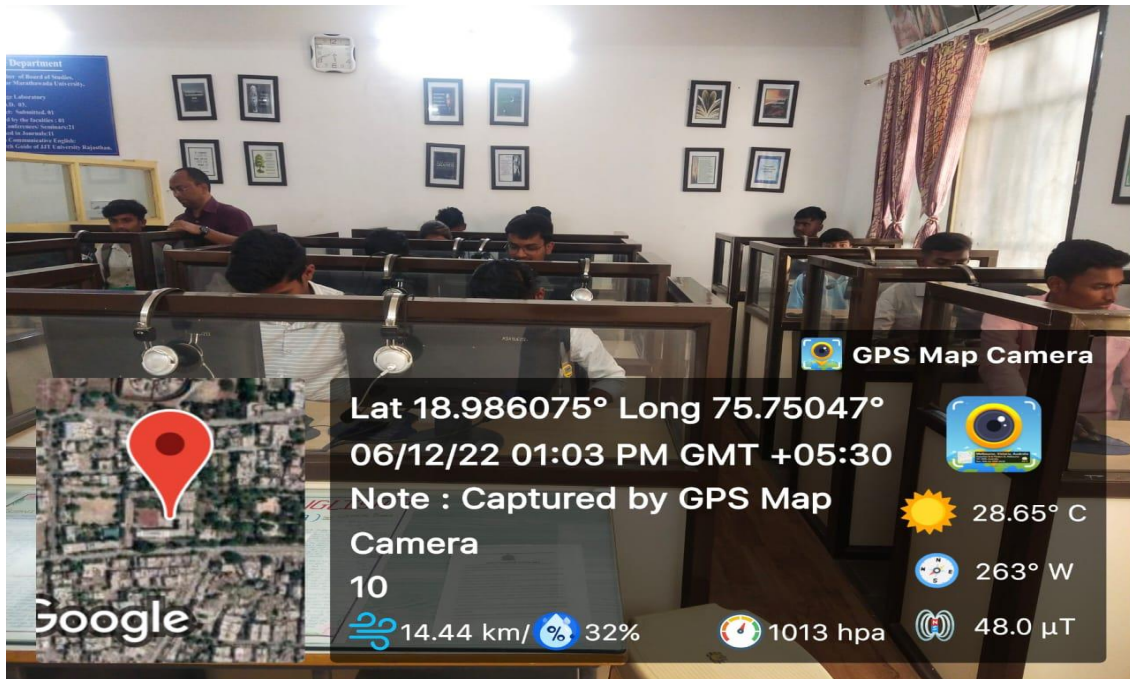
The problems encounter by the student:

- ❖ There is a lack of vocabulary.
- ❖ There is less scope for innovative research.
- ❖ Student just memorized the dialogue.
- ❖ Interaction is predominantly one sided as compared to class room.

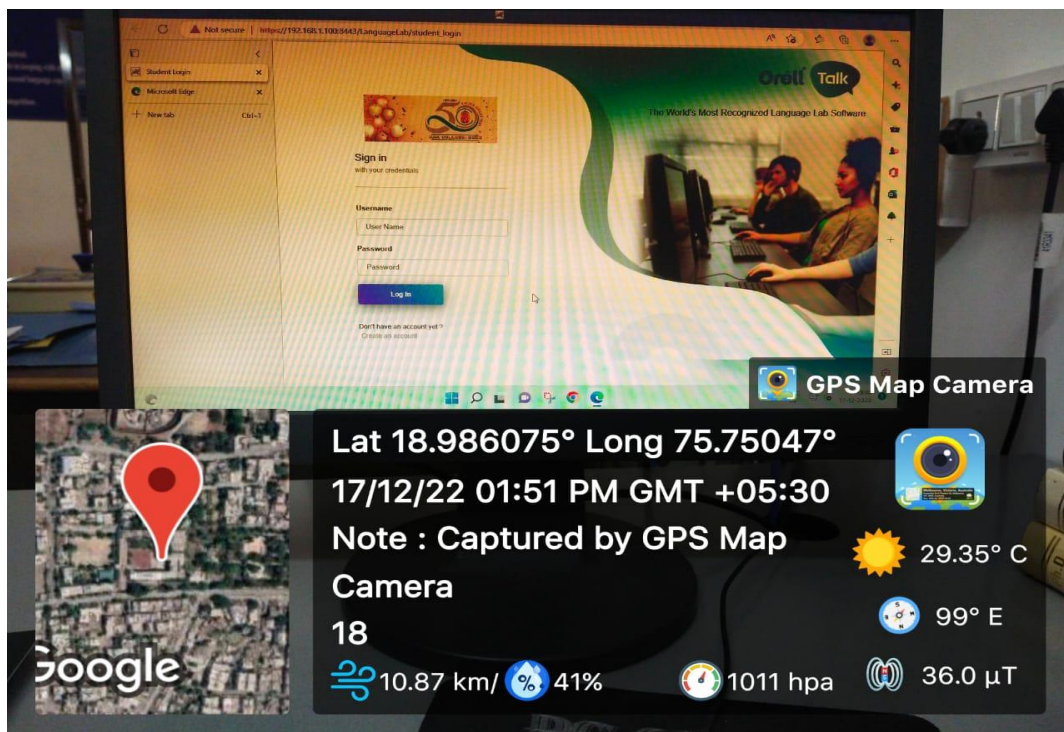
7) Resources Required :-

- ❖ Computers are very important part in the language laboratory
- ❖ latest technology
- ❖ head set
- ❖ Well furniture
- ❖ sufficient electricity
- ❖ internet facilities
- ❖ Software
- ❖ Study material.

Best Practice-1:- English Language laboratory; Photos



Students are using language software under the guidance of the teachers.



Language Lab software: Orell Talk

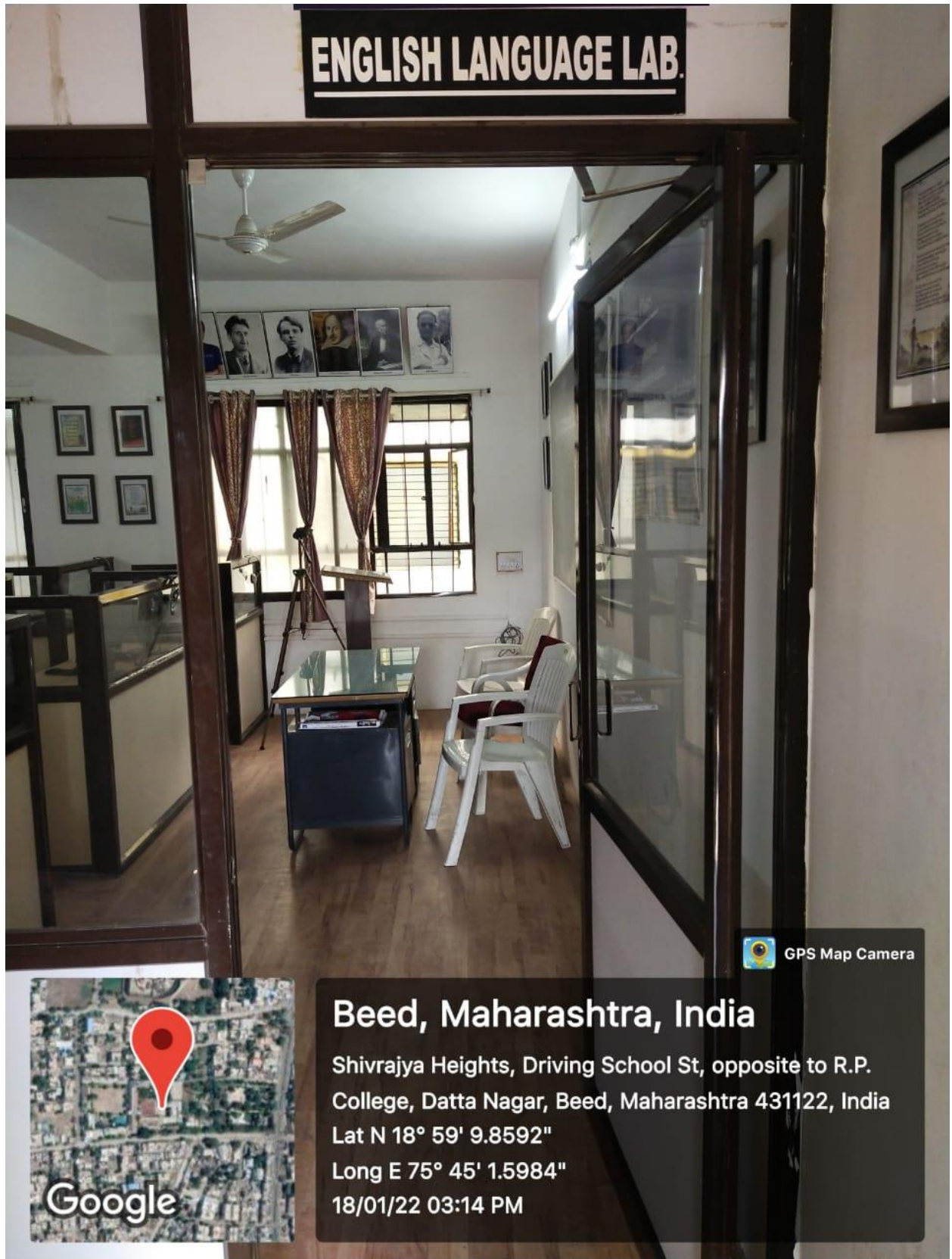
English Language Laboratory

List of Students enrolled for English Language Laboratory: 2021-22

Sr. No.	Name of the Student	Class
1	Adagale Keshav Govind	BA I
2	Doke Swati Bansi	BA I
3	Borwade Rohini Bhimrao	BSC II
4	Chepte Ganesh Hanumant	BSC II
5	Devkate Pawan Narayan	BA I
6	Gaikawad Ajay Kanta	BA I
7	Aghav Arti Arjun	BSC I
8	Gaikawad Sumit Balu	BA I
9	Gore Lakhan Ram	BA I
10	Maske Savari Ashok	BA I
11	Mane Keshav Bhagwan	BA I
12	Khade Diksha Machindra	BA I
13	Missal Vithal Ram	BA I
14	Khope Sneha Vijay	BA I
15	Mehetre Krushna Ganesh	BA I
16	Bande Pawan Vasant	BA I
17	Chavan Nutan Anurath	BA I
18	Amte Pallavi Bandu	BA I
19	Gotral Rahul Kailas	BA II
20	Fasale Ayodhya Ramchandra	BA II
21	Bhole Pratik Arunrao	BA II
22	Dhole Manish Ashok	BA II
23	Kate Rohit Kisan	BA II
24	Bahirwal Sudhakar Rajabhau	BSC I
25	Bhosale Pratik Satish	BSC I
26	Baglane Pawan Narayan	BSC I
27	Bhosale Mohini Ashruba	BSC I
28	Joshi Gaurav Kalidas	BSC II
29	Kadam Mayuri Vikas	BSC II
30	Jadhav Aditya Ankush	BSC II
31	Shinde Datta Bhimrao	BSC II
32	Thokal Vishal Babasaheb	BSC II
33	Vaidya Kalyani Prerak	BSC II

34	Shinde Meena Namdeo	BSC II
35	Todekar Rushikesh Muralidhar	BSC II
36	Tandale Tushar Madhukar	BSC II
37	Zade Avinash Shivaji	BSC II

Head, Dept. of English



GPS Map Camera



Beed, Maharashtra, India

Shivrajya Heights, Driving School St, opposite to R.P.
College, Datta Nagar, Beed, Maharashtra 431122, India
Lat N 18° 59' 9.8592"

Long E 75° 45' 1.5984"

18/01/22 03:14 PM



GPS Map Camera



Beed, Maharashtra, India

Shivrajya Heights, Driving School St, opposite to R.P. College,
Datta Nagar, Beed, Maharashtra 431122, India

Lat N 18° 59' 9.906"

Long E 75° 45' 1.6272"

18/01/22 03:11 PM



GPS Map Camera



Beed, Maharashtra, India

Shivrajya Heights, Driving School St, opposite to R.P. College,
Datta Nagar, Beed, Maharashtra 431122, India

Lat N 18° 59' 9.906"

Long E 75° 45' 1.6272"

18/01/22 03:11 PM



GPS Map Camera

Beed, Maharashtra, India

Shivrajya Heights, Driving School St, opposite to R.P. College,
Datta Nagar, Beed, Maharashtra 431122, India

Lat N 18° 59' 9.816"

Long E 75° 45' 1.62"

18/01/22 03:11 PM

Google



Highlights of the Department

- Dr. Tonde S.R. is a Member of Board of Studies, Dr. Babasaheb Ambedkar Marathwada University, Aurangabad.
- Well Equipped Language Laboratory
- Faculties completed Ph.D. 03.
- Minor Research project: Submitted: 01
- Research Papers Published by the faculties : 01
- Conferences/ Seminars: 21
- ...

GPS Map Camera



Beed, Maharashtra, India

Shivrajya Heights, Driving School St, opposite to R.P. College, Datta Nagar, Beed, Maharashtra 431122, India
Lat N 18° 59' 9.8916"
Long E 75° 45' 1.7712"
18/01/22 03:11 PM



Google



Cased
ge Lab
e

Orell Talk

LIVE VIRTUAL CLASS



The world's most
Recognized
Digital Language
Laboratory System

Acquiring intelligible English
with "Global Accent"



e-Exam

Online Examination System
with Instant Scores/Assessment



Career Lab

Career Oriented Courses,
Personality Development
& Etiquette Building

Live Classroom

Virtual Classes & Video
Conferencing with
Interactive Whiteboard

- ✓ User-friendly & Cost-effective
- ✓ CEFR Levels based Learning
- ✓ Choice of Cloud & Offline versions
- ✓ Teacher-led & Self-study options



www.orelltalk.com

18/01/22 03:07 PM

Long E 75° 45' 1.7568"

Lat N 18° 59' 9.8808"

Beed, Maharashtra, India
Shivajya Heights, Driving School St, opposite to R.P. College,
Data Nagar, Beed, Maharashtra 431122, India

24x7 Support
+91 2388-814-000



GPS Map Camera

OrellTalk

Shivaji N. Shinde

KSK Alias Kaku College of Arts, Science

Dashboard

Class View

Click Here to view Students

Things to do

view Study Materials

e-Assignments

Click Here to view Student Activities

Grade Summary

Sachin Yadav

Kiran George

Abul Khan

Phil Johns

Nikhil Kumar

David Rajan

Attendance

of Votes

Type here to search

28°C Sunny

15:09

18-01-2022

Google

Beed, Maharashtra, India

Shivrajya Heights, Driving School St, opposite to R.P. College,

Datta Nagar, Beed, Maharashtra 431122, India

Lat N 18° 59' 9.888"

Long E 75° 45' 1.7568"

18/01/22 03:09 PM

Best Practice – 2

1. Title of the practice: Remedial classes for slow learners

2. Objectives of the practice:

- To encourage the academically weaker students to identify and work on their weakness and help them to improve on their learning skills.
- To enhance the level of understanding of basic concepts in various subjects.
- To provide a stronger base for further academic progress.
- To strengthen their knowledge skills and attitude in such subjects.
- To improve academic skills of the students in various subjects.

3. The Context:

One of the objectives of our college is to provide good quality education to students from rural areas and to develop their interest in education. Beed district in Maharashtra is considered to be underdeveloped and backward in terms of development. Almost half of the students admitted to our college are from underprivileged society of rural areas. Due to the lack of good educational institutions in the rural areas, the students are found to be out of the mainstream of education. Also, due to the poor financial situation of the house, they have to contribute to the work of their parents every day. Therefore, these students could not take proper education to strengthen their knowledge skills and academic skills. Hence the college has decided to conduct remedial classes in all subjects for these students so as to bring them in a mainstream of the education.

4. The Practice:

At the beginning of every academic year, a test is conducted to identify the academically slow learners. Notice regarding remedial class is circulated among the students who are aspiring and willing to attend the remedial classes. Slow learners are also identified during the regular classes. Once the slow learners are identified, a curriculum based on some basic concepts is designed by the head and professors of the corresponding departments. Systematic schedule is prepared for the delivery of content of remedial coaching classes with the consent and permission of the principal of the college. Remedial classes are organised

during the two days of the week so as to meet the needs of the slow learners. Scheduled content of the remedial class is delivered during two months. One of the professors in the department is given the responsibility of identifying and resolving the problems of slow learners. During the course of remedial class, tests are conducted regularly to monitor the progress of slow learners and at the end of the remedial class course, an examination is conducted so as to observe the progress of the slow learners.

5. Evidences of Success:

Due to these remedial classes, the progress of the students can be monitored and the teachers get to know the weakness and the strengths of the students so that the teachers can plan properly for their overall development. These classes solidify the basic concepts of the students and they understand the main core concepts of each subject as well. This eventually resulted in progress in their final year results. Mainly, due to these classes, confidence was shown by these students to confront any exam.

6. Problems encountered and Resources Required:

In-charge professors were allotted an extra workload to conduct remedial classes but this turned out to be an excellent piece of devotion towards these students. To conduct these classes, the overall time table of the college had to be reshuffled. Time constraint is always there to conduct these classes as it requires quite more time so as to address the different problems of the students. More ICT based content and ICT tools are required to enhance the interests of the students in the concerned subjects.

Best Practice-2:- Remedial Classes for slow learners; Photos

Remedial Coaching Course: Syllabus

B.Sc. II Year Chemistry(2020-2021)

Total : 24 Hours , Three hour per Week

Unit - I Chemical compounds & their families 06 Hours

- Types of chemical compounds,
- Functional groups of organic compounds
- Different theories of Acids and Bases,

Unit – II Basic concepts in Organic Chemistry: 06 Hours


- Chemical Bonds & it's types ,
- Intermolecular forces & it's types,
- Mechanism of Chemical reactions,
- Hydrocarbons and types
- Heterocyclic compounds

Unit – III Spectroscopy; 06 Hours

- Brief account on spectroscopy,
- Types of spectroscopy,
- Elucidation of spectral data,

Unit – IV Laboratory skills in Performing Experiments: 06 Hours

- Operating Instruments,
- Handling Apparatus, glass wares etc.
- Preparation of Standards Solutions of different Normality, Molarity.


Head
Department of Chemistry
Mrs. K.S.K. College, Beed.

Remedial Coaching Course: Syllabus

B.Sc. I Year Chemistry (2021-2022)

Total: 24 Hours , Three hour per Week

Unit –I Mathematical Concepts: 06 Hours

- Basic mathematics used in Physical Chemistry
- Differentiation, Integration, etc.
- Logarithmic Rules;
- Internal conversions; numbers in points to powers.
- Practice of drawing various types of graphs in Physical Chemistry,

Unit – II Physical Properties and it's Units: 06 Hours

- Measurements of Physical Properties,
- Units of Physical Properties,
- Unit Conversions; SI units to CGS units,

Unit – III Laboratory skills in Performing Experiments: 06 Hours

- Operating Instruments,
- Handling Apparatus, glass wares etc.
- Preparation of Standards Solutions of different Normality, Molarity.

Unit – IV Atomic Structure: 06 Hours

- History of atom , Atomic orbital's,
- Quantum numbers, periodic properties
- Electronic configurations of the elements,

Head

Department of Chemistry
Mrs. K.S.K. College, Beed,


NSSR's
Mrs. Kesharbai Sonajirao Kshirsagar alias Kaku Arts, Science &
Commerce College, Beed.


Department of Chemistry

Time – Table for the Remedial Course; 2021-2022

B.Sc. I, II Year.

Day / Time	Class	Name of the Teacher	Hall No.
Monday 7.30 – 8.20 am	B. Sc. II Yr.	Dr.P.R.Khakre	50
Tuesday 7.30 – 8.20 am	B. Sc. II Yr.	Dr.P.R.Khakre	50
Wednesday 7.30 – 8.20 am	B. Sc. II Yr.	Dr.P.R.Khakre	50
Thursday 7.30 – 8.20 am	B. Sc. I Yr.	Dr.S.B.Maulage	50
Friday 7.30 – 8.20 am	B. Sc. I Yr.	Dr.S.B.Maulage	50
Saturday 7.30 – 8.20 am	B. Sc. I Yr.	Dr.S.B.Maulage	50


HOD – Chemistry
Department of Chemistry
Mrs. K.S.K. College Beed.


Principal

NSSR's

**Mrs. Kesharbai Sonajirao Kshirsagar alias Kaku Arts,
Science & Commerce College, Beed.**

Department of Chemistry

Remedial Coaching Course

B.Sc. II Year Chemistry (2021-2022)

Student list

Sr.no	Name of student
1.	Akhade Nikhil Vishnu
2.	Dake Trupti Indrajit
3.	Gaikwad Kartik Sundar
4.	Ghule Omkar Govindrao
5.	Kale Akshay Sadashiv
6.	Late Ajit Parmeswar
7.	Pathak Gajanan Girishkumar
8.	Phad Santosh Manik
9.	Rajpure Archana Bandu
10.	Sable Umesh Laxman
11.	Shinde Meena Namdev
12.	Shinde Vaibhavi Ashruba
13.	Sirsat Rohit Vilas
14.	Sonwane Mayur Bhaskar
15.	Survase Aishwarya Babasaheb
16.	Takale Omkar Hanuman
17.	Tandle Prachi Rajabhau
18.	Thokal Sakshi Santosh
19.	Thokal Vishal Babasaheb
20.	Thombre Rushikesh Murlidhar
21.	Vaidya Kalyani Prerak
22.	Varad Bhagyashree Santosh
23.	Waghmare Pormina Navnath
24.	Wahed Shaikh Malik Wahed Shaikh
25.	Wakhare Vaishnavi Vitthal

24.	Salve Saurabh Rajendra
25.	Takik vaishnavi Sham
26.	Yadav Jay Vijay



Head
Department of Chemistry
Mrs. K.S.K. College, Beed.



NSSR's

**Mrs. Kesharbai Sonajirao Kshirsagar alias Kaku Arts,
Science & Commerce College, Beed.**

Department of Chemistry

Remedial Coaching Course

B.Sc. I Year Chemistry (2021-2022)

Student list

Sr.no	Name of student
1.	Aghav Aarti Arjun
2.	Akhade Sagar Satish
3.	Barde Karan Kailas
4.	Chate Shubham Bappasaheb
5.	Dhas Tushar Khanderao
6.	Dhase Pavan Ankush
7.	Gadale Ankush Janak
8.	Garud Aniket Ganesh
9.	Ghuge Rutuja Balu
10.	Jadhav Pratik Balasaheb
11.	Joshi Kanchal Dilip
12.	Kale Nikhil Datta
13.	Kamble Rahul Sunil
14.	Kate Ganesh Bharat
15.	Kekan Gorakh Dhondiba
16.	Korde Nishant Balu
17.	Kute Nitin Babasaheb
18.	Magar Anand Sanjay
19.	Mane Pradnya Sham
20.	Munde Sunanda Ramesh
21.	Nagargoje Pratiksha janardhan
22.	Pawar Avinash Sanjay
23.	Rasal Pooja Vasant

Department of Chemistry
2021-22
Remedial Class Test – B.Sc.II year

Name :	Marks obtained =
Date:	Out of = 20

Multiple choice questions with one correct answer.

- The general formula for alkanes is
a) $C_n H_{2n}$ b) $C_n H_{2n+2}$ c) $C_n H_{2n-2}$ d) None of these
- Alkanes mainly show
a) Ionic elimination reactions b) Ionic formation reaction
c) Free radical elimination reactions d) Heat/Photochemical reactions
- Which one of the following bonds is strongest ?
a) $-C-C-$ b) $-C=C-$ c) $-C\equiv C-$ d) None of these
- Which one of the following is not found in free state ?
a) Au b) Na c) Cu d) Ag
- Alkali metals are generally prepared by
a) Reducing the corresponding oxides with Mg b) Reducing halides with hydrogen
c) Electrolytic reduction of their molten halides d) Reduction of metal oxides with carbon
- The decreasing order of reactivity of hydrogen atoms attached to different carbon atoms in alkanes is
a) Tertiary > Primary > Secondary b) Tertiary > Secondary > Primary
c) Primary > Secondary > Tertiary d) Secondary > Primary > Tertiary
- Iodination of alkanes is best carried out in the presence of
a) H_2O b) HIO_3 c) C_6H_6 d) NH_4SH
- Alkali metal have how many electron in their outer most shell
a) 0 b) 1 c) 2 d) 1 or 2
- Which of the following is the most abundant alkali metal (in combined state) in nature ?
a) Li b) Na c) Cs d) K
- In the series, ethane, ethylene and acetylene, the C-H bond energy is
a) Least in acetylene b) Greater in ethane c) Greatest in acetylene d) Equal in all
- Electrochemical process is employed to extract....
a) Mercury b) Cadmium c) sodium d) silver
- The maximum ease of abstraction of a hydrogen atom by a chlorine atom is shown by
a) $(CH_3)_4C$ b) $(CH_3)_3CH$ c) $C_6H_5CH_3$ d) $CH_2=CHCH_3$
- The reaction $RX + 2 Na + RX \rightarrow R-R + 2NaX$ is called
a) Fittig reaction b) Sabatier and Senderson's reaction c) Wurtz's reaction d) Wurtz fittig reaction
- A mixture of C_2H_6 , C_2H_4 and C_2H_2 is passed through ammonical $AgNO_3$ solution. The gases which remain unreacted are
a) C_2H_6 and C_2H_4 b) C_2H_6 and C_2H_2 c) C_2H_4 and C_2H_2 d) None of these
- Element of 1st group give colour in Bunsen burner due to
a) Low ionisation potential b) Low melting point c) High reactivity d) One electron in their outermost shell
- The ashes of plants contain alkali metal, 90% of which is
a) Li b) K c) Cs d) Ca
- Formation of alkanes by the action of Zn on alkyl iodide is called
a) Frankland reaction b) Cannizzaro reaction c) Kolbe's reaction d) Wurtz reaction
- Which of the following electronic configuration represents an alkali metal ?
a) ns^1 b) ns^2 c) ns^2np^1 d) ns^1np^2
- Alkali metal are
a) Reducing agents b) Oxidising agents c) Both reducing and oxidising agents d) Complexing agents
- The minimum first ionisation energy is shown by which of the following electronic configuration
a) $1s^2 2s^1$ b) $1s^2 2s^2 2p^6 3s^2 3p^2$ c) $1s^2 2s^2 2p^6 3s^1$ d) $1s^2 2s^2 2p^6$

[Signature]
Head

Department of Chemistry
M.S.K. College, Beed.



Remedial Coaching Classes (2021-22)

Introduction:

The department of chemistry has been suggested by Hon. Principal to conduct the remedial classes for the slow learners in the subject of chemistry. Therefore the chemistry department has decided to run the remedial classes for weak students/ slow learners in the subject of chemistry.

Selection procedure of Students:

At the beginning of every academic year, the weak students or slow learners are identified through oral discussion and class performance. These students are considered for the enrollment of remedial class.

Syllabus Frame:

As per the need of students, the syllabus for the remedial classes has been framed and the same is distributed into two parts for B.Sc. I and B.Sc. II year students. 24 hours have been allotted to complete the concerned syllabus of each class (B.sc. I & II Year). The said course syllabus for each class is expected to complete within two months therefore for each class 03 hours per week have been allotted.

Student's progress:

After the completion of syllabus of each class successfully, students progress is evaluated by conducting a test of 20 marks comprising 20 multiple choice questions based on the syllabus framed for the remedial class. The result of the students (mark list of tests) is displayed on the notice board after five days from the date of test conducted.

Besides, the faculty of chemistry is always motivating the students to participate in various activities of the college and continuously encourage and cooperate them by counseling.


Head
Department of Chemistry
Mrs. K.S.K. College, Beed.

SAMPLE

NSSR'S

Mrs. Kesharbai Sonajirao Kshirsagar alias Kaku Arts, Science and Commerce
College, Beed.

Department of Chemistry

Analysis report of Student's Feedback for Remedial Classes

For the academic year – 2021-22

Sr. No	Particulars	% of Very Good	% of Good	% of Satisfactory	% of Unsatisfactory
1.	Remedial class course content				
2.	Relevance of remedial course content with main courses				
3.	Learning resources for the remedial class (Library, ICT etc.)				
4.	Helpful in improving the subject knowledge				
5.	Various skills inculcated through course				
6.	Assignments/Evaluation transparency				
7.	Interest generated by the teacher				
8.	Extent of coverage of course during prescribed period				

Head
Department of Chemistry
Mrs. K.S.K. College, Beed.



SAMPLE

NSSR'S

Mrs. Kesharbai Sonajirao Kshirsagar alias Kaku Arts, Science and Commerce College, Beed.

Student's Feedback for Remedial Class

Name of the student: _____ Class: _____

Department: _____ Year of Study: _____

Make a tick mark in the appropriate cell :

Sr. No	Particulars	Very Good	Good	Satisfactory	Unsatisfactory
1.	Remedial class course content				
2.	Relevance of remedial course content with main courses				
3.	Learning resources for the remedial class (Library, ICT etc.)				
4.	Ease of the content for conceptual understanding				
5.	Various skills inculcated through course				
6.	Assignments/Evaluation transparency				
7.	Interest generated by the teacher				
8.	Extent of coverage of course during prescribed period				

Suggestions for further improvement:

Signature

Head

Department of Chemistry
Mrs. K.S.K. College, Beed,

Head

Department of Chemistry
Mrs. K.S.K. College, Beed.

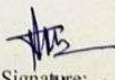
Remedial Class Test Mark Sheet -2019-20

Department of Chemistry

Class: B.Sc. I Year

Sr. No.	Name of Students	Mark Obtained
1	Ade Ravina Baburao	09
2	Abire Anamika Bhagwat	08
3	Bhoskar Y. S	10
4	Chavan Ganesh Sudam	07
5	Phere Sancheti Sandipam	04
6	Gade Pami Lalasabeb	07
7	Gaikwad Pami Balasabeb	10
8	Hange Sakshi Santosh	07
9	Inkar Sonket Madhukar	09
10	Jagtap Amol Hamidas	10
11	Kadam Dadasabeb Anag	08
12	Kale Y. K	08
13	Mandve Prajakt M	08
14	Netke Sachin Sanjivrao	08
15	Orhal Vishal Kailas	10
16	Patil Ganesh Kailas	06
17	Pathod Sonal Shesherao	09
18	Sonap Vaibhav. Bhagchand	08
19	Syed Sahil Qaisar	08
20	Tule Vijay Ankuab	07
21	Narade Prayonka Uddhan	08
22	Wagh Arpit Abhimanyu	07
23	Ware Harshada Dilip	05
24	Wagh Sunil Kishan	10
25	Yele Ashwini Bapu	09
26		

Name of Teacher: Dr. Khakre P. R.


 Signature: _____
 Head
 Department of Chemistry
 Mrs. K.S.K. College, Beed.



Scanned with CamScanner

Remedial Class Test Mark Sheet -2019-20

Department of Chemistry

Class: B.Sc. II Year

Sr. No.	Name of Students	Mark Obtained
1	Agwan Vijay Babasaheb	08
2	Bade Dhyaneshwar Laxman	08
3	Banekar Amol Ashok	10
4	Chavan Abhijit Mahadev	09
5	Chavan Pratik Deepak	08
6	Cholse Gajanan Ashok	07
7	Dake Rohini Namdeo	09
8	Dolas Kiran Mangabhai	08
9	Gaukwaad Shital Balasaheb	09
10	Gholap Ganesh Bhaskar	09
11	Ingle Varsha Santosh	09
12	Jagtap Tushar Anil	09
13	Kale Amol Chandu	10
14	Kale Swati Balasaheb	09
15	Mane Ashwini Gokul.	08
16	Munde Abhishek Dilip	09
17	Nismal Aditi Kishor	09
18	Parse Mahadev Babasaheb	08
19	Potdar Pravin Vitthal	09
20	Raut Marayan Sundar	09
21	Sonap Shrutti Madhukar	08
22	Thite Shahaji Vijaynath	08
23	Thosar Akshay Mahadev	10
24	Ugale Prashant Shivram	08
25	Waghre Krushna Ramu	08
26	Zodge Pratik Kishanrao	08

Name of Teacher: Dr. S. B. Maulage.


Signature:



Scanned with CamScanner

Department of Chemistry
Remedial Progress Test
2021-22

Name :	Marks obtained =
Date:	Out of = 20

Multiple choice questions with one correct answer.

- Isotones have same number of
 - Protons
 - Electrons
 - Neutrons
 - Protons and Neutrons
- The percentage of ionic character in HCl molecule is approximately
 - 71%
 - 100%
 - 12%
 - 17%
- The density of nucleus is of the order of
 - 10^{14} g/cm^3
 - 10^{12} g/cm^3
 - 10^{10} g/cm^3
 - 10^8 g/cm^3
- Isotopes have same number of
 - Protons
 - Electrons
 - Neutrons
 - None of these
- Isobars have same
 - Atomic number
 - Atomic mass
 - Both A and B
 - None of these
- The value of Planck's constant h is numerically equal to
 - ergs sec
 - J sec
 - Both A and B
 - Neither A or B
- Maximum covalency of an element is limited to
 - 8
 - 9
 - 6
 - 5
- The energy of an electron in an atom is
 - Negative
 - Zero
 - Positive
 - None of the above
- Covalent compounds are soluble in
 - Polar solvents
 - Non-polar solvents
 - Water
 - None of the above
- The favourable condition (s) for the formation of an ionic compound is/are
 - Lower value of ionisation energy of an atom for cation formation
 - Higher value of electron affinity of an atom for anion formation
 - Higher lattice energy of the resultant ionic compound
 - All of the above
- Which is true about ionic bonds
 - Directional nature
 - Stereo Isomerism
 - High melting and boiling points
 - Noncrystalline nature
- Ionic compounds, in general, possess both
 - High melting points, boiling points and non-directional bonds
 - High melting points and low-boiling points
 - Directional nature of bonds and low boiling points
 - High solubilities in polar and non-polar solvents
- Molten sodium chloride conducts electricity due to the presence of
 - Free electrons
 - Free ions
 - Free molecules
 - Atoms of sodium and chlorine
- Most predominantly ionic compound will be formed by the combination of the groups
 - 1 and 17
 - 2 and 16
 - 3 and 15
 - 1 and 18
- Of the following solvents, the one most likely to dissolve ionic compounds is
 - Carbon tetrachloride
 - Methanol
 - Water
 - Benzene
- The pair of elements which form a compound with maximum ionic character is
 - Na and O
 - Na and I
 - Cs and I
 - Cs and F
- Covalent bond is formed by the,
 - Sharing of electrons between two atoms
 - Transference of electrons from one atom to the other
 - Sharing of two electrons between two atoms, when the electrons are contributed by one atom only
 - None of these
- Tetravalency of carbon is best explained by the concept of
 - Resonance
 - Hybridization
 - Electron delocalization
 - None of these
- What types of bonds are generally formed between like atoms ?
 - Ionic
 - Covalent
 - Coordinate
 - All of these
- The maximum covalency is generally equal to
 - The number of unpaired s-electrons
 - The number of paired p-electrons
 - The number of unpaired s-and p-electrons
 - The actual number of s-and p-electrons present in the valence shell



Scanned with CamScanner


 Head
 Department of Chemistry
 Mrs. K.S.K. College, Beeg

Date: 25/10/2021

To,

**The Principal,
Mrs. K. S. K. Arts, Sci. & Comm. College,
Beed.**

Sub: Permission for remedial coaching classes – Chemistry department.

Respected Madam,

Faculty of chemistry department has identified the weak students / slow learners in subject through oral discussion and classroom performance. These students are needed to be considered for the enrollment of remedial classes. Department of chemistry, therefore is willing to conduct remedial coaching during 01/11/2021 to 31/01/2022 for identified slow learners from B.Sc. I & II year classes so as to bring them in the mainstream of average learners in the chemistry subject.

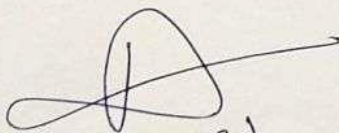
Therefore you are requested to give permission to start the remedial coaching by the department of Chemistry.

Thanking you.

Yours Faithfully



Dr. P.R. Khakre
Head, Dept. of Chemistry
Department of Chemistry
Mrs. K.S.K. College, Beed.



26.10.21